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back forward main menu toc study guide toc - $\text{H}_2\text{O}(l) \rightarrow \text{H}_2(g) + \frac{1}{2}\text{O}_2(g)$ $\Delta H = 296.36 \text{ kJ}$ $\text{SO}_2(g) \rightarrow \text{S}(\text{monoclinic}) + \text{O}_2(g)$ calculate the enthalpy change for enthalpy change for the reaction Study Guide TOC Textbook

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